

Name: Key Class: _____ Date: _____

Chemical Nomenclature Worksheet
Ionic compounds, covalent molecules, and acids...oh my!

Write chemical formulas for problems 1-20 and names for problems 21-40.

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|----------------------------|---|--|---------------------------------|
| 1. sulfur dioxide | <u>SO₂</u> | 21. NiF ₂ | <u>nickel(II) fluoride</u> |
| 2. ammonium phosphate | <u>(NH₄)₃PO₄</u> | 22. AlCl ₃ | <u>aluminum chloride</u> |
| 3. lithium hydroxide | <u>LiOH</u> | 23. N ₂ O ₄ | <u>dinitrogen tetroxide</u> |
| 4. sodium sulfate | <u>Na₂SO₄</u> | 24. AgC ₂ H ₃ O ₂ | <u>silver acetate</u> |
| 5. perbromic acid | <u>HBrO₄</u> | 25. H ₂ SO ₃ | <u>sulfurous acid</u> |
| 6. chloric acid | <u>HClO₃</u> | 26. HCN | <u>hydrocyanic acid</u> |
| 7. iron(II) chromate | <u>FeCrO₄</u> | 27. SiO ₂ | <u>silicon dioxide</u> |
| 8. silicon dioxide | <u>SiO₂</u> | 28. CO | <u>carbon monoxide</u> |
| 9. carbon tetrachloride | <u>CCl₄</u> | 29. BaSO ₃ | <u>barium sulfite</u> |
| 10. hydrochloric acid | <u>HCl</u> | 30. PF ₅ | <u>phosphorus pentafluoride</u> |
| 11. nickel(II) sulfite | <u>NiSO₃</u> | 31. Al ₂ O ₃ | <u>aluminum oxide</u> |
| 12. diphosphorus pentoxide | <u>P₂O₅</u> | 32. FeO | <u>iron(II) oxide</u> |
| 13. copper(II) sulfate | <u>CuSO₄</u> | 33. HBr | <u>hydrobromic acid</u> |
| 14. silver iodide | <u>AgI</u> | 34. H ₂ C ₂ O ₄ | <u>oxalic acid</u> |
| 15. acetic acid | <u>HC₂H₃O₂</u> | 35. CO ₂ | <u>carbon dioxide</u> |
| 16. arsenic acid | <u>H₃AsO₄</u> | 36. NH ₄ I | <u>ammonium iodide</u> |
| 17. platinum(IV) cyanide | <u>Pt(CN)₄</u> | 37. HNO ₃ | <u>nitric acid</u> |
| 18. carbon disulfide | <u>CS₂</u> | 38. HNO ₂ | <u>nitrous acid</u> |
| 19. gold(III) chloride | <u>AuCl₃</u> | 39. MgBr ₂ | <u>magnesium bromide</u> |
| 20. sulfurous acid | <u>H₂SO₃</u> | 40. SnS ₂ | <u>tin(IV) sulfide</u> |

Name: Key Class: _____ Date: _____

Chemical Nomenclature Worksheet #2
More ionics, covalents, and acids...oh my!

Write chemical formulas for problems 1-20 and names for problems 21-40.

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|-----------------------------|----------------------------------|--------------------|-----------------------------------|
| 1. thiocyanic acid | <u>HSCN</u> | 21. H_2CrO_4 | <u>chromic acid</u> |
| 2. copper(I) carbonate | <u>Cu_2CO_3</u> | 22. IF_7 | <u>iodine heptafluoride</u> |
| 3. ammonium sulfite | <u>$(NH_4)_2SO_3$</u> | 23. $Sn_3(PO_4)_4$ | <u>tin(IV) phosphate</u> |
| 4. hydroiodic acid | <u>HI</u> | 24. P_3N_5 | <u>triphosphorus pentanitride</u> |
| 5. diphosphorus pentoxide | <u>P_2O_5</u> | 25. ClO_2 | <u>chlorine dioxide</u> |
| 6. antimony(III) chloride | <u>$SbCl_3$</u> | 26. $NaCN$ | <u>sodium cyanide</u> |
| 7. magnesium hydroxide | <u>$Mg(OH)_2$</u> | 27. CoF_3 | <u>cobalt(III) fluoride</u> |
| 8. silver hypochlorite | <u>$AgClO$</u> | 28. XeF_4 | <u>xenon tetrafluoride</u> |
| 9. ammonia | <u>NH_3</u> | 29. NI_3 | <u>nitrogen triiodide</u> |
| 10. iodic acid | <u>HIO_3</u> | 30. KF | <u>potassium fluoride</u> |
| 11. sulfur trioxide | <u>SO_3</u> | 31. $H_2C_2O_4$ | <u>oxalic acid</u> |
| 12. zinc oxide | <u>ZnO</u> | 32. $HClO$ | <u>hypochlorous acid</u> |
| 13. silver cyanide | <u>$AgCN$</u> | 33. HCl | <u>hydrochloric acid</u> |
| 14. phosphorus pentabromide | <u>PBr_5</u> | 34. $NaMnO_4$ | <u>sodium permanganate</u> |
| 15. barium bisulfite | <u>$Ba(HSO_3)_2$</u> | 35. $HMnO_4$ | <u>permanganic acid</u> |
| 16. calcium bicarbonate | <u>$Ca(HCO_3)_2$</u> | 36. PbO | <u>lead(II) oxide</u> |
| 17. lead(IV) chlorite | <u>$Pb(ClO_2)_4$</u> | 37. KOH | <u>potassium hydroxide</u> |
| 18. beryllium arsenate | <u>Be_3As_2</u> | 38. CaH_2 | <u>calcium hydride</u> |
| 19. aluminum perchlorate | <u>$Al(ClO_4)_3$</u> | 39. As_2O_5 | <u>diarsenic pentoxide</u> |
| 20. nickel(II) iodide | <u>NiI_2</u> | 40. HF | <u>hydrofluoric acid</u> |