

Name: Key Class: \_\_\_\_\_ Date: \_\_\_\_\_

**Chemical Nomenclature Worksheet – Writing Formulas**  
**Ionic compounds, covalent molecules, and acids...oh my!**

Write chemical formulas.

1. sulfur dioxide	<u>SO<sub>2</sub></u>	21. thiocyanic acid	<u>HSCN</u>
2. ammonium phosphate	<u>(NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub></u>	22. copper(I) carbonate	<u>Cu<sub>2</sub>CO<sub>3</sub></u>
3. lithium hydroxide	<u>LiOH</u>	23. ammonium sulfite	<u>(NH<sub>4</sub>)<sub>2</sub>SO<sub>3</sub></u>
4. sodium sulfate	<u>Na<sub>2</sub>SO<sub>4</sub></u>	24. hydroiodic acid	<u>HI</u>
5. perbromic acid	<u>HBrO<sub>4</sub></u>	25. diphosphorus pentoxide	<u>P<sub>2</sub>O<sub>5</sub></u>
6. chloric acid	<u>HClO<sub>3</sub></u>	26. antimony(III) chloride	<u>SbCl<sub>3</sub></u>
7. iron(II) chromate	<u>FeCrO<sub>4</sub></u>	27. magnesium hydroxide	<u>Mg(OH)<sub>2</sub></u>
8. silicon dioxide	<u>SiO<sub>2</sub></u>	28. silver hypochlorite	<u>AgClO</u>
9. carbon tetrachloride	<u>CCl<sub>4</sub></u>	29. ammonia	<u>NH<sub>3</sub></u>
10. hydrochloric acid	<u>HCl</u>	30. iodic acid	<u>HIO<sub>3</sub></u>
11. nickel(II) sulfite	<u>NiSO<sub>3</sub></u>	31. sulfur trioxide	<u>SO<sub>3</sub></u>
12. diphosphorus pentoxide	<u>P<sub>2</sub>O<sub>5</sub></u>	32. zinc oxide	<u>ZnO</u>
13. copper(II) sulfate	<u>CuSO<sub>4</sub></u>	33. silver cyanide	<u>AgCN</u>
14. silver iodide	<u>AgI</u>	34. phosphorus pentabromide	<u>PBr<sub>5</sub></u>
15. acetic acid	<u>HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub></u>	35. barium bisulfite	<u>Ba(HSO<sub>3</sub>)<sub>2</sub></u>
16. arsenic acid	<u>H<sub>3</sub>AsO<sub>4</sub></u>	36. calcium bicarbonate	<u>Ca(HCO<sub>3</sub>)<sub>2</sub></u>
17. platinum(IV) cyanide	<u>Pt(CN)<sub>4</sub></u>	37. lead(IV) chlorite	<u>Pb(ClO<sub>2</sub>)<sub>4</sub></u>
18. carbon disulfide	<u>CS<sub>2</sub></u>	38. beryllium arsenate	<u>Be<sub>3</sub>(AsO<sub>4</sub>)<sub>2</sub></u>
19. gold(III) chloride	<u>AuCl<sub>3</sub></u>	39. aluminum perchlorate	<u>Al(ClO<sub>4</sub>)<sub>3</sub></u>
20. sulfurous acid	<u>H<sub>2</sub>SO<sub>3</sub></u>	40. nickel(II) iodide	<u>NiI<sub>2</sub></u>

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**Chemical Nomenclature Worksheet – Writing Names**  
More ionics, covalents, and acids...oh my!

Write chemical names.

- |                                       |                                 |                                      |                                   |
|---------------------------------------|---------------------------------|--------------------------------------|-----------------------------------|
| 1. $\text{NiF}_2$                     | <u>nickel(II) fluoride</u>      | 21. $\text{H}_2\text{CrO}_4$         | <u>chromic acid</u>               |
| 2. $\text{AlCl}_3$                    | <u>aluminum chloride</u>        | 22. $\text{IF}_7$                    | <u>iodine heptafluoride</u>       |
| 3. $\text{N}_2\text{O}_4$             | <u>dinitrogen tetroxide</u>     | 23. $\text{Sn}_3(\text{PO}_4)_4$     | <u>tin(IV) phosphate</u>          |
| 4. $\text{AgC}_2\text{H}_3\text{O}_2$ | <u>silver acetate</u>           | 24. $\text{P}_3\text{N}_5$           | <u>triphosphorus pentanitride</u> |
| 5. $\text{H}_2\text{SO}_3$            | <u>sulfurous acid</u>           | 25. $\text{ClO}_2$                   | <u>chlorine dioxide</u>           |
| 6. $\text{HCN}$                       | <u>hydrocyanic acid</u>         | 26. $\text{NaCN}$                    | <u>sodium cyanide</u>             |
| 7. $\text{SiO}_2$                     | <u>silicon dioxide</u>          | 27. $\text{CoF}_3$                   | <u>cobalt(III) fluoride</u>       |
| 8. $\text{CO}$                        | <u>carbon monoxide</u>          | 28. $\text{XeF}_4$                   | <u>xenon tetrafluoride</u>        |
| 9. $\text{BaSO}_3$                    | <u>barium sulfite</u>           | 29. $\text{NI}_3$                    | <u>nitrogen triiodide</u>         |
| 10. $\text{PF}_5$                     | <u>phosphorus pentafluoride</u> | 30. $\text{KF}$                      | <u>potassium fluoride</u>         |
| 11. $\text{Al}_2\text{O}_3$           | <u>aluminum oxide</u>           | 31. $\text{H}_2\text{C}_2\text{O}_4$ | <u>oxalic acid</u>                |
| 12. $\text{FeO}$                      | <u>iron(II) oxide</u>           | 32. $\text{HClO}$                    | <u>hypochlorous acid</u>          |
| 13. $\text{HBr}$                      | <u>hydrobromic acid</u>         | 33. $\text{HCl}$                     | <u>hydrochloric acid</u>          |
| 14. $\text{H}_2\text{C}_2\text{O}_4$  | <u>oxalic acid</u>              | 34. $\text{NaMnO}_4$                 | <u>sodium permanganate</u>        |
| 15. $\text{CO}_2$                     | <u>carbon dioxide</u>           | 35. $\text{HMnO}_4$                  | <u>permanganic acid</u>           |
| 16. $\text{NH}_4\text{I}$             | <u>ammonium iodide</u>          | 36. $\text{PbO}$                     | <u>lead(II) oxide</u>             |
| 17. $\text{HNO}_3$                    | <u>nitric acid</u>              | 37. $\text{KOH}$                     | <u>potassium hydroxide</u>        |
| 18. $\text{HNO}_2$                    | <u>nitrous acid</u>             | 38. $\text{CaH}_2$                   | <u>calcium hydride</u>            |
| 19. $\text{MgBr}_2$                   | <u>magnesium bromide</u>        | 39. $\text{As}_2\text{O}_5$          | <u>diarsenic pentoxide</u>        |
| 20. $\text{SnS}_2$                    | <u>tin(IV) sulfide</u>          | 40. $\text{HF}$                      | <u>hydrofluoric acid</u>          |